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## PERIODIC- TEST [ Solutions ]

**Full marks:- 20** [ Set-A ] **Time:- 45 Minutes**

Very Short Answer Type Questions:- 1X8 = 8

1. Define non-aqueous solution with example.

2. How many solute particles are in heptanary solutions?

3. What is the molality of pure water?

4. What happen if more solute is added in super saturated solution?

5. Why molality is preferred for experimental data in place of Molarity?

6. Give an example of solution of liquid in gas.

7. What happen to Saturated solution if temperature increases?

8. Define unsaturated solution.

Numerical Problems:- 3X4 = 12

1. 30 g Urea (NH2CONH2) is dissolved in 90 g of water to prepare its

 aqueous solution. Determine the mole fraction of each component and

 molality of the solution.

2. Calculate the molality of 2.5g of ethanoic acid (CH3COOH) in 75g of

 benzene.

3. 20 g of NaOH is present in 90 g of water to prepare it’s aqueous solution

 What would be the molality, molarity and mole fraction of each component

 of the solution? The density of the solution is 1.1 g/ml.

4. What would be the molality and mole fraction of an aqueous solution

 of glucose of concentration 5M? Density of the solution is 1.1 g/ml.



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